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been investigated.

Surface tension and surface
thermodynamic properties of

Thus the energy consumed is
at least the thermodynamic
surface energy W_s arising

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?, $W_s = 4 \pi r_0^2 \gamma$. At the
critical equilibrium point,
the closing stress $2\gamma / r_0$
due to the surface tension
acting at the boundary of
the cavity is equal to the

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hydrostatic tension ρh , and
hence $r_0 = 2\gamma / \rho h$.

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advertenties weer te geven.

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Surface tension is an important factor in the phenomenon of capillarity.

Surface tension has the dimension of force per unit length, or of energy per unit area. The two are

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equivalent, but when referring to energy per unit of area, it is common to use the term surface energy, which is a more general term in the sense that it applies also to solids.

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studies of surface-tension,
 $\sigma = ?$. Temperature
dependence can be determined
by differentiating with
respect to temperature. •
Thus, temperature dependence
of surface tension leads to

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the specific surface
entropy. • Temperature
dependence of γ predicting
by a semi-empirical equation
(van der Waals and
Guggenheim: where $\gamma_0 = \gamma$ at
 $T = 0 \text{ K}$, T

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In the preceding chapter,
the surface tension and the

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related thermodynamic quantities of two-component two-phase systems were discussed in detail. In this Science

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vergelijkbare tools om uw
winkelervaring te
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advertentities were te given.

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From 251 data points
examined the average
relative deviation (ARD) of

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the correlated and
calculated densities and
surface tension from the
experimental ones was found
to be 1.63% and 2.46%.

Besides, some surface
thermodynamic functions such
as the surface entropy (S S)

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and surface enthalpy (H_S) of studied liquids were also computed via our method. The ARD (in %) were found to be equal to 3.94 and 2.44, respectively.

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