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## Stoichiometry Practice Test Answers

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(in grams) of ammonia produced when...

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Answers: Moles and Stoichiometry Practice Problems 1) How  
many moles of sodium atoms correspond to  $1.56 \times 10^{21}$  atoms  
of sodium?  $1.56 \times 10^{21}$  atoms Na  $\times$  1 mol Na =  $2.59 \times 10^{-3}$   
mol Na  $236.022 \times 10$  atoms Na 2) Determine the mass in

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grams of each of the following: a. 1.35 mol of Fe 1.35 mol Fe x 55.845 g Fe = 75.4 g Fe 1 mol Fe b. 24.5 mol O

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Answers to Stoichiometry Practice Free Response Test: 1. 30.40 g. N X 1 mol N/14.01g N = 2.17 moles N. 69.60 g O X 1 mol. O /16.00 g O = 4.35 moles O. 4.35/2.17. = 2.00 2.17/2.17 = 1.00 So O is 2X more present than N. Empirical formula must be NO. 2. Mass.

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## Stoichiometry Practice Test With Answers

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## Chapter 12 Stoichiometry Test Answer Key

Practice Test Ch 3 Stoichiometry Name Per. Remember it is a MC test, use the answers ... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: 2 NH<sub>3</sub> + 5 O<sub>2</sub> → 2 NO<sub>2</sub> + 3 H<sub>2</sub>O ... 7. c First you must realize this is a limiting reactant problem. Found: 3 Jan 2020 | Rating: 89/100

## Stoichiometry Practice Problems Level 1 Answers

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20 Then do some stoichiometry using "easy math" 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of CO<sub>2</sub> = 44 g, and 2 moles of H<sub>2</sub>O which is 36 g for a total of 80 g.

4. d Balance: C<sub>3</sub>H<sub>8</sub> + 5O<sub>2</sub> → 3CO<sub>2</sub> + 4H<sub>2</sub>O

5. d Balance: 2KClO<sub>3</sub> → 2KCl + 3O<sub>2</sub>

## Practice Test Ch 3 Stoichiometry Name Per

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Stoichiometry Practice Test - Answer Key. Back to the other Stoichiometry Practice Tests and other General Chemistry Practice Tests. The formation of NH<sub>3</sub> from N<sub>2</sub> and H<sub>2</sub> occurs in 85.0% yield. How many grams of ammonia would be experimentally obtained when 12 g of H<sub>2</sub> reacts with 20g of N<sub>2</sub>?

## Stoichiometry Practice Test - Answer Key - SarahChem

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Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

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Determine the amount (in moles) of a product from a given amount of one reactant. Determine the amount (in moles) of a product from a given amount of one reactant. If you're seeing this message, it means we're having trouble loading external resources on our website.

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