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Solutions Class 12 Notes Chemistry Chapter 2. 1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

Solutions Class 12 Notes Chemistry Chapter 2 - Learn CBSE 1. Solution Chemistry Notes. 2. Solutions • Basic definitions – Solution: homogenous mixture – Solute: substance that is dissolved – Solvent: substance that contains the solute in a solution • Process of Dissolving: – Molecules of the solvent “ pull apart ” the solute molecules, and distributes them within the solvent • How to tell the solute from the solvent – The substance that is in the greater amount is the solvent – The substance that keeps its phase is the solvent.

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Chemistry Notes for class 12 Chapter 2 Solutions Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

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Solutions Solutions: Solutions are the homogeneous mixtures of two or more than two components. Binary solution: A solution having two components is called a binary solution. Components of a binary solution. It includes solute and solvent. When ...

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assignment with complete solutions. Physicswallah Typed ...

## 01. Solutions - PhysicsWallah

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Chemistry Solution Notes. Classification. Solutions which  
contain two components in it are called Binary Solutions.  
Substances which are used to prepare a solution are called

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as Components. The component that is present in the largest quantity is known as Solvent. Solvent determines the physical state in which solution exists.

Chemistry Solution Notes CBSE Class 12 pdf (Important ...  
Types of solutions, expression of concentration of solutions of solids in liquids, solubility of gases in liquids, solid solutions, colligative properties - relative lowering of vapour pressure, Raoult's law, elevation of boiling point, depression of freezing point, osmotic pressure, determination of molecular masses using colligative properties, abnormal molecular mass, Van't Hoff factor.

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CBSE Class 12 Chemistry Notes, Short Key Notes, Practice ...  
Revision Notes on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of

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which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute.

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