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bonds do not give all of the
atoms in the molecule eight
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bonds to other atoms and two

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lone pairs. The shape of
this ...

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geometry in which 3 atoms

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are not in a straight line.
This geometry occurs when
the central atoms contain 4
electron groups (2 bonding
and 2 nonbonding) or 3
electron groups (2 bonding
and 1 nonbonding) The bond
that results when two

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nonmetals combine in a chemical reaction.

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In this chapter, we describe

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how electrons are arranged in atoms and how the spatial arrangements of electrons are related to their energies. We also explain how knowing the arrangement of electrons in an atom enables chemists to predict

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and explain the chemistry of an element.

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Electrons move in circular orbits . around . the nucleus . at . fixed.

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energy levels. Electrons are never between energy levels or energy shells. An electron must have just the right amount of energy to jump from one level to another. A ... Chapter 13
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Atoms and Ions Chemical Reactivity The electrons in the outer orbit (valence electrons)

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Chapter 5: Electrons in Atoms. STUDY. PLAY. 3×10^8 m/s. What is the speed of light? meters. What unit must wavelength be in? $E=h\nu$. Equation to solve for energy. $\nu = E/h$. Equation for frequency. Joules. In what

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units is energy measured.

$6.626 \times 10^{-34} \text{ J}\cdot\text{s}$. Planck's constant. Speed of light formula. $c = \lambda \times \nu$.

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Worksheet of the Atom
Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to

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explain why objects change color when heated.

Chapter 5 Electrons In Atoms

Answer Key

Chapter 5: Electrons in Atoms
5.1 Wave-Particle Duality/Electromagnetic

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Spectrum/Relationship of Wavelength, Frequency and Speed of light
5.2 Bohr's Model of the Atom/Quantum Mechanical Model of the Atom
5.3 Electron Arrangement & Valence Electrons

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Chapter 9: Electrons in Atoms and the Periodic Table
Ch 9 Page 1 . Visible light -the form of electromagnetic radiation that we can "see"
Microwaves, radiowaves, x

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-rays, ultraviolet radiation, infrared radiation -the forms of electromagnetic radiation that we can't see.

**Chapter 9: Electrons in
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Resource Manager Objectives
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and You will express the
arrangements of electrons in
atoms through orbital
notations, electron Analysis
Answers will vary. Students

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will determine that observations typically rely heavily upon sight, ...

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Atoms Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and

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frequency are inversely related; in other words, as one quantity increases, the ...

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