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In 1924, Louis de Broglie proposed a new speculative hypothesis that electrons and other particles of matter can behave like waves. Today, this idea is known as de Broglie's hypothesis of matter waves. In 1926, De Broglie's hypothesis, together with Bohr's early quantum theory, led to the development of a new theory of wave quantum mechanics to describe the physics of atoms and subatomic particles.

De Broglie's Matter Waves - University Physics Volume 3

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Louis de Broglie - Wikipedia Louis-Victor-Pierre-Raymond, 7e duc de Broglie was a French physicist who made groundbreaking contributions to quantum theory. In his 1924 PhD thesis he Solving the measurement problem: de Broglie-Bohm loses...1 Feb 2008 A common attitude amongst those that take the de Broglie-Bohm interpretation ..Perhaps the most obvious - and the most principled - defence is ...

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De Broglie's phase wave and periodic phenomenon. De Broglie's thesis started from the hypothesis, "that to each portion of energy with a proper mass  $m_0$  one may associate a periodic phenomenon of the frequency  $\nu_0$ , such that one finds:  $h\nu_0 = m_0 c^2$ . The frequency  $\nu_0$  is to be measured, of course, in the rest frame of the energy packet ...

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de Broglie equation states that a matter can act as waves much like light and radiation which also behave as waves and particles. The equation further explains that a beam of electrons can also be diffracted just like a beam of light. In essence, the de Broglie equation helps us understand the idea of matter having a wavelength.

de Broglie Equation - Derivation, de Broglie Wavelength ...

Abstract: Introductory course on quantum and atomic physics including optics and statistical physics. Objective: A basic introduction to quantum and atomic physics, including basics of optics and equilibrium statistical physics.

Course Catalogue - ETH Z

When considering de Broglie wave phenomena of larger and more complex objects than atoms, fullerenes come to mind as suitable candidates. After their discovery<sup>9</sup> and the subsequent invention of efficient mass-production methods<sup>10</sup>, they became easily available. In our experiment (see Fig. 1) we use commercial, 99.5% pure, C

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De Broglie waves account for the appearance of subatomic particles at conventionally unexpected sites because their waves penetrate barriers much as sound passes through walls. Thus a heavy atomic nucleus occasionally can eject a piece of itself in a process called alpha decay. The piece of nucleus (alpha particle) has insufficient energy as a particle to overcome the force barrier surrounding the nucleus; but as a wave it can leak through the barrier—that is, it has a finite probability ...

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