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Chemquest 33 Answers

Chemquest 33 Limiting Reactants Answers the “ limiting reactant ” and oxygen is the excess reactant. For each mole of C₃H₈ five moles of O₂ are required, so for 12.5 moles of C₃H₈, the number of moles of O₂ needed are $(12.5)(5) = 62.5$ moles.

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Chemquest 33 Answers (Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) 12 moles. For every mole of propane that combusts 3 moles of CO₂ are produced, so the number of moles of CO₂ that can be produced when 4 moles of propane combusts = $4(3) = 12$.
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Chemquest 33 Limiting Reactants Answers the “ limiting reactant ” and oxygen is the excess reactant. For each mole of C₃H₈ five moles of O₂ are required, so for 12.5 moles of C₃H₈, the number of moles of O₂ needed are $(12.5)(5) = 62.5$ moles.

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c) $3.1 \times 10^{-34} + 2.2 \times 10^{-33} =$ first, change exponents so both are either -34 or -33 ... sheet. Any substance listed on the periodic table. For example, phosphorus or nitrogen.

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ChemQuest 33 Name _____,uTg.eTlnT F.uT'r \$\$ Date Hour Information:

Determining if a Bond is Polar In general the greater the difference in electronegativity between two bonding atoms, the greater the polarity of the bond. A general rule of thumb is that if the difference in electronegativity is less than 0.5 then the bond is considered nonpolar.

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ChemQuest 33. ChemQuest 33. Name: _____ . Date:

_____. Hour: _____. Information: Limiting Reactant. Again consider the combustion of propane: $C_3H_8 + 5 O_2 \rightarrow 3 CO_2 + 4 H_2O$. If you had 10 moles of propane to burn, you would need 50 moles of oxygen according to the ratio in the balanced equation.

ChemQuest 33 - Webs

ChemQuest 10 Name: _____ Date: _____ Hour: _____ Information: Bohr ' s Solar System Model of the Atom All the planets are attracted to the sun by gravity. The reason that the Earth doesn ' t just float out into space is because it is constantly attracted to the sun. Similarly, the moon is attracted to the Earth by

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Critical Thinking Questions: Bohr ' s reasoning

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moles of C₃H₈, the number of moles of O₂ needed are $(12.5)(5) = 62.5$ moles.
Since we have more than 62.5 moles (according to the question we have Page 4/26
Chemquest 33 Limiting Reactants Answers ChemQuest 33 Name T tlt'cym
,uTg.eTInT F.uT'r \$\$ Date Hour

Chemquest 33 Limiting Reactants Answers

PLEASE NOTE: If you have a question about these answers, it is your responsibility
to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys -
ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 -
HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of

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JI ChemQuest 34 Name: Date Hour::F-F Fluorine Gas Molar mass = 38.0 g c.t- Ql Chlorine Gas Molar mass ::Br Br Bromine Liquid Molar mass : I-I' Iodine Solid Molar mass : Information: Some Molecules and Their States Critical Thinking Questions
What type of force exists between two F₂ molecules-- {ispersion, dipolar, or hydrogen bonds?,".eh"s#n ?(€l'5Feq-ece\$'1 What t5.pe of force exists between ...

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Since we have more than 62.5 moles (according to the question we have Page 4/26

Chemquest 33 Limiting Reactants Answers ChemQuest 33 Name T tl'tcym

,uTg.eTlnT F.uT'r \$\$ Date Hour

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ChemQuest 3 Name: _____ Date: _____ Hour: _____ Information: Changes in Matter

Books are made of matter. You are made of matter. “ Matter ” is a fancy word for the “ stuff ” of which all objects are made. Every day, matter is changed in different ways.

For example, paper can be changed in many ways—it can be torn, folded, or burned. ...

Information : Changes in Matter - Science! - Home

A certain element exists as three different isotopes. 24.1% of all isotopes have a mass of 75.23 amu, 48.7% have a mass of 74.61 amu, and 27.2% have a mass of 75.20 amu.

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Chemquest 29 Answers 24 ChemQuest 29 Monc LcW s Stnuctunes tr I Name Date:

Hour: Introduction Questions Draw the Lewis structures for the following molecules.

Part c is done for you. a) H₂O (oxygen is central) b) SO₂ c) N₂ It t-t (a [Jrt @ N=N @ @ f\ @ 2 In question 1 above you should see three different kinds of bonds of bonds as best you can.

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