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Moles, Examples
& Practice
Problems**

Converting Units

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Factors Unit

Conversion the

Easy Way

(Dimensional

Analysis)

Converting Units

with Conversion

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Converting Units

using Multiple

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How to Convert

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~~Convert Between~~

~~Grams and Moles~~

Unit Conversion

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Understanding

Conversion

Factors Multiple

Conversion With

Factors (Part 2)

Dimensional

Analysis:

Converting Units

with 1 or 2

Conversion

Factors Metric

Unit Prefix

Conversions: How

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The quantity in
the problem is
multiplied by
one or more
“conversion

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factors," in
which the
numerator is
equal to the
denominator.

Since the
numerator and
denominator of
the conversion
factor are
equal,
multiplying by
the conversion
factor is like

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multiplying by 1
and thus does
not change the
value of the
original
quantity.

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Factor Problems
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chemistry units

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quiz!.

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did well on the
units and

conversions

quiz. If you
have trouble
with any

specific types
of problems, try
looking at a

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worked example
problem to
review the
concepts and see
how to proceed.
Remember to
check your work
to make sure an
answer makes
sense.

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molecule. For example, you might be given moles or grams of C_2H_6 and be asked to find the number of H atoms in that sample. In problems like this, you can use the fact that there are 6 atoms of H in 1

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mlcl of C₂H₆ as
a conversion
factor.

Factors Practice

~~Chemical
Conversions and
Problems~~

Converting
Between Units
with Conversion
Factors. A
conversion
factor is a
factor used to

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convert one unit
of measurement
into another. A
simple
conversion
factor can be
used to convert
meters into
centimeters, or
a more complex
one can be used
to convert miles
per hour into
meters per

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Problems For

conversions

within the

metric system,

you must

memorize the

conversion (for

example: 1000 mL

= 1 L, or 1000 g

= 1 kg should be

memorized)

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Remember that

metric

conversions are

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Multiple

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Conversions.

Sometimes you
will have to
perform more

than one
conversion to
obtain the
desired unit.

For example,
suppose you want
to convert 54.7
km into
millimeters. We
will set up a

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series of
conversion
factors so that
each conversion
factor produces
the next unit in
the sequence.

~~2.7: Solving
Multi-step
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Chart — Density,
Volume, Grams to
...

Now let's
practice a few
unit conversion
problems using
conversion

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factors.

Practice Problem

#1 Convert 45.0
inches to

centimeters,

given that 1

inch equals 2.54
centimeters.

~~Conversion~~

~~Factor in~~

~~Chemistry:~~

~~Definition,~~

~~Formula...~~

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Practice:

Convert units
(metrics) This
is the currently
selected item.

Metric units of
mass review (g
and kg) Metric
units of length
review (mm, cm,
m, & km) Metric
units of volume
review (L and
mL) U.S.

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customary and
metric units.

Next lesson.

Converting

metric units

word problems.

~~Convert units~~

~~(metrics)~~

~~(practice) |~~

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Valence

Electrons

Practice

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easy cross
reference. The
Overview section
summarizes the
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includes a
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terms, a summary
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concepts, and a
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exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in, Multiple Choice are two sets of

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questions that
include every
concept and
numerical
exercise
introduced in
the chapter and
the Skills
section provides
developed
exercises to
apply the new
concepts in the
chapter to

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particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and

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multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of

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followed by
discussions on
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