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**Ionic and Metallic
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Metals Pt III What Are

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Introduction to Ionic
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and Ionic Compounds

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Section 7.3 – Bonding
in Metals. The
valence electrons of
metal atoms can be
modeled as a sea of
electrons. Metallic
bonds consist of the
attraction of the free-
floating valence

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electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

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... Bonding

1. Chapter 7 "Ionic
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2. Section 7.1 - Ions

OBJECTIVES:

–Determine the
number of valence
electrons in an atom
of a representative
element. 3. Section

7.1 - Ions

OBJECTIVES:

–Explain how the

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octet rule applies to
atoms of metallic and
nonmetallic elements.

4.
Answer Key

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Compounds and

Metals Section 77.1.1

Figure 7.1 As carbon

dioxide dis-solves in

ocean water,

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carbonate ions are produced. Coral polyps capture these carbonate ions, producing crystals of calcium carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

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Compounds and
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Chapter 7 Ionic and Metallic Bonding Answer Key Multiple Choice

Identify the choice that best completes the statement or answers the question.

_____ 1. What are the valence electrons of an atom? a. The electrons that an atom gains when it becomes an anion b.

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The electrons that remain closest to the nucleus c. The electrons in the highest occupied energy level d. The electrons that an atom loses when it ...

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valence electrons.

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And Metallic
electron dot

structures. octet rule.
halide ions. electrons
in the highest

occupied energy level
of an element's....

diagrams that show
valence electrons in
the atoms of an

elemen.... in forming
compounds, atoms
tend to achieve the
electron confi....

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Bonding. BONDING
AND

INTERACTIONS. 7.1

Ions. Essential Unders

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Bonding
Answer Key

standing Ions form
when atoms gain or
lose valence
electrons, becoming
electrically charged.

Lesson Summary.

Valence

Electrons Valence

electrons are the
electrons in the
outermost occupied
energy level and are
involved in ion
formation. For a

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representative
element, the group
number equals the ...
Answer Key

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OBJECTIVES: –

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number of valence
electrons in an atom

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octet rule Ionic and
covalent bonding
Bond polarity and
electronegativity
Formal charge
Chapter 8 Energetics
of bond formation and
dissociation Bond
formation: $A + B \rightarrow AB$ If
the reaction is
energetically

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favorable, AB is lower
in energy than A + B.
Heat is released .

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