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 $0.110 \text{ g} \times 1 \text{ washer} = 11.0 \text{ g}$ (assuming 100 washers is exact) $100. \text{ g} \times 1 \text{ washer} = 909 \text{ washers}$
2. $500. \text{ g} \times 1 \text{ cork} = 1.63 \text{ g} \times 306.7 = 307 \text{ corks}$
 $500. \text{ g} \times 1 \text{ stopper} = 4.31 \text{ g} \times 116 \text{ stoppers} = 1 \text{ kg}$ (1000 g) of corks contains (1000 g

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Chapter 6 Equilibrium Chemistry 211 the reaction we monitor the mass of Ca^{2+} remaining in solution and the mass of CaCO_3 that precipitates, the result looks something like Figure 6.1. At the start of the reaction the mass of Ca^{2+} decreases and the mass of CaCO_3 increases. Eventually the reaction reaches a point after which there

Chapter 6

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