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matthewmotamedi. Hydrocarbon Compounds. Terms in this set (66) hydrocarbons ...

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22CHAPTER Study Guide Key Concepts 22.1 Hydrocarbons

- Because carbon has four valence electrons, carbon atoms always form four covalent bonds.
- The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches.
- Molecules of hydrocarbons, such as alkanes, are nonpolar molecules.

22.2 Unsaturated Hydrocarbons

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A controlled process by which hydrocarbons are broken down or rearranged into smaller, more useful molecules natural gas This is typically made up of 80% methane, 10% ethane, 4% propane, and 2% butane

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718 Chapter 22 22CHAPTER Study Guide Key Concepts 22.1 Hydrocarbons • Because carbon has four valence electrons, carbon atoms always form four covalent bonds. • The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches.

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Chapter 22: Substituted Hydrocarbons and Their Reactions.

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784 Formic acid. BIGIdeaThe substitution of different functional groups for hydrogen atoms in hydrocarbons results in a diverse group of organic compounds. 22.1 Alkyl Halides and Aryl Halides. MAINIdeaA halogen atom can replace a hydrogen atom in some hydrocarbons.

Chapter 22: Substituted Hydrocarbons and Their Reactions

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Chapter 22 - Hydrocarbon Compounds - 22.1 Hydrocarbons

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Hydrocarbons are organic molecules that only contain two elements: carbon and hydrogen. Common hydrocarbons

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include methane, propane, and butane. Hydrocarbons are the primary constituents of fossil...

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