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K_w pH pOH pK_a

pK_b H^+ OH^-

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Answer Key

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Questions 1.

What is true of
all chemical
processes? 2.

How can you
distinguish a
system and its
surroundings? 3.

What are the
required
characteristics

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of an isolated system? 4.
Explain the
difference

between
exothermic and
endothermic
processes. 5.

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Key Chapter 17

...

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ochemistry 187
10. Complete the
enthalpy diagram
for the
combustion of
natural gas. Use
the
thermochemical
equation in the
first paragraph
on page 517 as a
guide. SECTION
17.3 HEAT IN

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Chapter 17 Thermodynamics
CHANGES OF STATE
(pages 520–526)

This section explains heat transfers that occur during melting, freezing, boiling, and condensing.

*SECTION 17.1 THE
FLOW OF ENERGY
HEAT AND WORK*

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(pages 505–510)

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Chapter 17 Thermodynamics (pages 505-510) This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat. ~ Energy Transformations (page 505)

1. What area of study in chemistry is

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concerned with
the heat
transfers that
occur during
chemical reactio
ns?thermochemist
ry.

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thermochemistry?

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4-11: Heat,
energy,
calorie/joule
conversions

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Notes, Slides
12-13: Endo vs.
Exo Demo:

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ammonium nitrate
+ water
(endothermic)
AND sodium
hydroxide +
water
(exothermic)
Chapter 17
Notes, Slides
15-17: Heat ...

*McLaughlin,
Kimberly /
Thermochemistry
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mochemistry
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heats of

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formation are
used to
calculate the
enthalpy change
for the reaction
of carbon
monoxide and
oxygen.

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Diagrams How

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Matching Match

each item with
the correct
statement below.

- a. calorimeter
- d. enthalpy b.
- calorie e.
- specific heat c.
- joule f. heat
- capacity _____ 1.
- quantity of heat

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needed to raise the temperature of 1 g of water by 1 °C 2. _____ SI unit of energy 3. _____ quantity of heat needed to change the temperature of 1 g of a ...

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ENERGY—HEAT AND

WORK (pages

505–510) This

section explains

the relationship

between energy

and heat, and

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distinguishes
between heat
capacity and
specific heat.

Energy
Transformations
(page 505)

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