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1 . 5.3 × × 10 3 C 3 . (a) reduction; (b) oxidation; (c) oxidation; (d) reduction 5 . (a) F 2 + Ca 2F ? + Ca 2+ ; F 2 + Ca 2F ? + Ca 2+ ; (b) Considerations include: cost of the materials used in the battery, toxicity of the various components (what constitutes proper disposal), should it be a primary or secondary battery, energy requirements (the “size” of the battery/how long ...

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Rhonda_FrazierTEACHER. Glencoe Chemistry Chapter 17. reversible recation. law of chemical equilibrium. equilibrium constant expression. Keq> 1. chemical reaction that can occur in borh the forward and the r…. particular ratio of reactant and product concentrations has a…. keq= [C]^c [D]^d/ [A]^a [B]^b.

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17-7 K and the extent of reaction K reflects a particular ratio of product concentrations to reactant concentrations for a reaction. A small value for K indicates that the reaction yields little product before reaching equilibrium. The reaction favors the reactants. K therefore indicates the extent of a reaction, i.e., how far a reaction proceeds towards the products at a given

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17.6 Corrosion – Chemistry
Chapter Outline. 17.1 Review of Redox Chemistry. 17.2 Galvanic Cells. 17.3 Electrode and Cell Potentials. 17.4 Potential, Free Energy, and Equilibrium. 17.5 Batteries and Fuel Cells. 17.6 Corrosion. 17.7 Electrolysis. Another chapter in this text introduced the chemistry of reduction-oxidation (redox) reactions.

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CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS
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