

Chapter 10 States Of Matter Test

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CHAPTER 10 REVIEW States of Matter SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. a. When a substance in a closed system undergoes a phase change and the system reaches equilibrium, (a) the two opposing changes occur at equal rates. (b) there are no more phase changes. (c) one phase change predominates.

10 States of Matter - Ms. Agostine's Chemistry Page

Modern Chemistry. Chapter 10. States of Matter. Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties. Kinetic-Molecular Theory of Gases 1- Gases consist of large numbers of tiny particles that are far apart relative to their size. 2- Collisions between gas particles and between particles and their container walls are elastic.

Modern Chemistry Chapter 10 States of Matter

Almost all substances can be classified into three states of matter - solids, liquids and gases. Each state has different properties. Heating and cooling a substance can cause it to change state.

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Chapter 10 Review States Of Matter Answers

CHAPTER 10 States Matter - phillipsmathandphysics.com Chapter 10 - states of matter 1. Chapter 10. States of Matter[] Kinetic Molecular Theory: the idea that particles of matter are always in motion; this applies to all states of matter[] Used to explain properties of matter in terms of energy 2.

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided Modern chemistry chapter 10 review answers states of matter section 1. 1. Match description on the right to the correct crystal type on the left Modern chemistry chapter 10 review answers states of matter section 1. b. Ionic crystal (a) has mobile electrons in the crystal c. covalent ...

Modern Chemistry Chapter 10 Review Answers States Of

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CHEMISTRY NOTES - Chapter 10 States of Matter Goals : To gain an understanding of : 1.. The kinetic theory of matter. 2. Avogadro's hypothesis. 3. The nature of liquids, solids and gases. 4. Phase changes. NOTES: The kinetic theory of matter states that particles which make up all types of matter are in constant motion. CHEMISTRY NOTES Chapter 10 States of Matter