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One of the most common titrations  
performed in a Chemistry lab is an acid-  
base titration. In the Initial Investigation,

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you will be assigned an acid solution to titrate with a solution of the strong base sodium hydroxide, NaOH. The concentration of the NaOH solution is given and you will determine the concentration of the acid solution.

Investigating Acid-Base Titrations -

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## Vernier

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Investigating Acid-Base Titrations. A  
titration is a laboratory process used to  
determine the volume of a solution needed  
to react with a given amount of another

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solution One of the most common titrations performed in a Chemistry . lab is an acid-base titration.

## Investigation 14 Investigating Acid-Base Titrations

In this lab, acid-base titrations will be performed in order to investigate the

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relationship between pH, concentration, and the structure of acids and bases. Students will carry out multiple acid-base titrations and will use a pH meter to monitor the pH of the resulting solutions. Students will then construct titration curves.

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## Ward's® AP Chemistry Investigation 14: Acid-Base ...

When looking at the graph of an acid base titration, the equivalence point is where the amount of acid and base is equal. The half-equivalence point is where the amount of base is equal to half the amount of acid in the solution. The curves for the

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titration of different strengths of acids and bases differ. A buffer solution is one that resists changes in pH.

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best completes the statement or answers the question. 1) The conjugate base of  $\text{HSO}_4^-$  is A)  $\text{H}_2\text{SO}_4$  B)  $\text{SO}_4^{2-}$  C)  $\text{H}_3\text{SO}_4^+$  D)  $\text{HSO}_4^+$  E)  $\text{OH}^-$  ... Advanced Placement Chemistry: 1997 Free Response (and Answers)

A.P. Chemistry Practice Test: Ch. 14,

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## Acids and Bases

When solving for pH using weak acid solutions, use the 5% rule. Divide the calculated value of  $H^+$  by the initial concentration of the weak acid (Ex. HF) and multiply it by 100, if its equal to or less than 5, the approximation can be assumed to be correct. Calculate  $H^+$  and

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pH.

## Calculating the pH of weak acid solutions - AP Chemistry

teacher, in the beaker. Repeat Steps 6-14,  
using acid. Testing the effect of base on  
other materials 17. Clean the 50 mL  
beaker labeled basic. Place 20 mL of test

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solution, obtained from your teacher, in the beaker. Repeat Steps 6–14, substituting base for acid. 18. If time permits, repeat Steps 16–17 for as many materials as you can. 19.

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Chemistry outline for 'Chapter 14 - Acids and Bases', from the Zumdahl's Chemistry, 5th Edition Textbook. These AP Chemistry notes will cover the key topics discussed in this chapter.

Chapter 14 - Acids and Bases |  
CourseNotes

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To find these values, we use the equation  $(K_b)(K_a) = K_w$  and  $K_w$  is equal to  $1.0 \times 10^{-14}$ . So, to solve for the  $K_b$  of a conjugate base, you divide  $K_w$  by the  $K_a$  of the weak acid, or for a...

Acid-Base Properties of Salts - AP  
Chemistry

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The main purpose of the lab is to perform different experiments and the importance of the topic is that it supports the essay's thesis statement. The question we're trying to answer is "How is pH used to determine if a solution is acidic or basic?" My hypotheses for this experiment is if the solution is an acid, then it will have a

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lower pH. If the solution is a base, it will have a ...

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