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The sum of the protons and neutrons in the nucleus of an atom.  $MASS\# = \#$

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protons(atomic number) + #neutrons.

#NEUTRONS = Mass # - #Protons(atomic number)  
#Protons = Mass # - #Neutrons.

ISOTOPEs: Are atoms of the same element that have different numbers of. neutrons and different mass numbers.

Chapter 4: Atomic Structure Notes

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2) All atoms of a given element are identical in size, mass, and chemical properties. 3) Atoms of a specific element different from

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those of another element. 4) Different atoms combine in simple whole-number ratios to form compounds. 5) A chemical reaction, atoms are separated combined or rearranged.

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